

Complete Table 3. Write the charge for both the cation and the anion in each formula.

| Table 3- Ion Charges | | |
|---|----------------------------|---------------------|
| | metal cation charge | anion charge |
| CuF | | |
| TiCl ₂ | | |
| CuSO ₄ | | |
| MnO | | |
| FeSe | | |
| PbI ₂ | | |
| Cr ₂ O ₃ | | |
| PdS | | |
| FeSO ₃ | | |
| Ti(NO ₃) ₃ | | |
| OsO ₄ | | |
| SnBr ₂ | | |
| UF ₆ | | |
| Pd(NO ₃) ₂ | | |
| Ag ₂ SO ₄ | | |
| Zn ₃ (PO ₄) ₂ | | |

Complete Table 4. Use solubility rules to determine if the compound is soluble in water.

| Table 4- Solubility of Ionic Compounds | | |
|---|---------------------|----------------------|
| | soluble (aq) | insoluble (s) |
| CuF | | |
| TiCl ₂ | | |
| CuSO ₄ | | |
| FeS | | |
| PbI ₂ | | |
| FeSO ₄ | | |
| Ti(NO ₃) ₂ | | |
| SnBr ₂ | | |
| Pd(NO ₃) ₂ | | |
| Ag ₂ SO ₄ | | |
| Zn ₃ (PO ₄) ₂ | | |